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range of the nuclear force allows them to attract only protons very close to them, as shown in Figure 22-3. So as the number of protons in a nucleus increases, the electrostatic force between protons increases faster than the nuclear force. More neutrons are required to increase the nuclear force and stabilize the nucleus.

CHAPTER 22 Nuclear Chemistry

Table 22.1 and example 22.1, problems 22.1-2 Beta emission (occurs when an atom has too many neutrons, or not enough protons). In the nucleus a neutron splits into an electron and a proton ($n \rightarrow p + e^{-}$) Since this electron originated from the nucleus, it is a beta particle.

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fusion, controlled nuclear reactions, radiation. Section 22.1 Define nucleon, isotope, nuclide, and nuclear reaction. Summarize the differences between nuclear reactions and chemical reactions. Section 22.2 Define radioactivity and radionuclide. Write balanced equations for nuclear reactions,

CHAPTER 22. NUCLEAR CHEMISTRY - Creighton University

Chapter 22: Nuclear Chemistry Section 22-1: The Nucleus • Atomic nuclei= protons and neutrons (together are nucleons) o Nuclide= an atom—identified by # of protons/neutrons in nucleus Mass Defect and Nuclear Stability • Mass defect= difference between mass of an atom and sum of the

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masses of protons/neutrons/electrons o Caused by conversion of mass to energy when nucleus forms Nuclear ...

Chapter 22 Notes - Chapter 22 Nuclear Chemistry Section 22 ...

1 CK-12 Physics Concepts - Intermediate Answer Key Chapter 22: The Nucleus 22.1 Isotopes Review Questions 1. What do different isotopes of a given element have in common? How are they different? 2. What is the element name of the atom represented by? 3. How many protons and how many neutrons does the nucleus in problem #2 have?

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