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Chapter 21 Nuclear Chemistry Section

Nuclear chemistry is the study of reactions that involve changes in

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nuclear structure. The chapter on atoms, molecules, and ions introduced the basic idea of nuclear structure, that the nucleus of an atom is composed of protons and, with the exception of ${}^1_1\text{H}$, neutrons.

21.1 Nuclear Structure and Stability - Chemistry

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Nuclear Chemistry Section 1 The

Nucleus Section 2 Radioactive Decay

Section 3 Nuclear Radiation Section 4

Nuclear Fission and Nuclear Fusion

CHAPTER 21. ... nuclear binding and

decay. 642 Chapter 21. C A B Number of

protons Number of neutrons 0 0 10 10

20 30 40 $n/p = 1.5:1$ $n/p = 1:1$ 50 60 70

80 90 100 20 30 40 50 60 70 80 90 100

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110 120 130

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Nuclear Chemistry Nuclear Transformations • Rutherford in 1919 performed the first nuclear transformation. • The transmutations are sometimes represented by listing in

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order, the target nucleus, the bombarding particle, the ejecting particle and the product nucleus. • The above equation becomes: ${}_{8}^{14}\text{N} + {}_{2}^{4}\text{He} \rightarrow {}_{7}^{14}\text{N} + {}_{1}^{1}\text{H}$

Chapter 21 Nuclear Chemistry

PDF Chapter 21 Nuclear Chemistry Notes (answers) Chapter 21: Nuclear

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Chemistry 21.1: The Nature of Nuclear Reactions
Nucleons: - the particles that make up a nucleus of an atom (protons, ${}^1_1\text{p}^+$ or ${}^1_1\text{H}$) and neutrons, ${}^1_0\text{n}$).
Isotopes: - atoms that have different mass number but the same atomic number or number of protons.

Chapter 21 Nuclear Chemistry

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Review Answers

Chapter 21 Nuclear Chemistry . Section 21.1 Types of Radioactivity. Objectives: Analyze Common Sources of Background Radiation, Compare and Contrast Alpha, Beta and Gamma Radiation, Apply the concept of Half-Life of a Radioactive Element. In a chemical reaction, what is the main subatomic particle involved?

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Chapter 21 Nuclear Chemistry

Modern Chemistry 171 Nuclearchemistry

CHAPTER 21 REVIEW Nuclear Chemistry

SECTION 2 SHORT ANSWER Answer the following questions in the space

provided. 1. _____ The nuclear equation is an example of an equation that represents (a) alpha emission. (b) beta

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emission. (c) positron emission. (d) electron capture. 2.

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Chemistry End of Chapter Exercises.
Write the balanced nuclear equation for
the production of the following
transuranium elements: (a)
berkelium-244, made by the reaction of

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Am-241 and He-4 (b) fermium-254, made by the reaction of Pu-239 with a large number of neutrons (c) lawrencium-257, made by the reaction of Cf-250 and B-11

21.4 Transmutation and Nuclear Energy - Chemistry

CHAPTER 21 REVIEW Nuclear Chemistry.

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CHAPTER 21 REVIEW Nuclear Chemistry
SECTION 4 SHORT ANSWER Answer the
following questions in the space
provided. 1. Match each of the following
statements with the process(es) to
which they apply, using one of the
choices below: (1) fission only (3) both
fission and fusion

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21.2 Nuclear Reactions and Energy.

Objectives: Compare and Contrast Nuclear fission and Fusion, Demonstrate Equations that Represent the Changes that Occur During Radioactive Decay, Trace the Operation and Structure of a Nuclear Reactor.

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A nuclear fuel. A fissionable isotope must be present in large enough

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quantities to sustain a controlled chain reaction. The radioactive isotope is contained in tubes called fuel rods. A moderator. A moderator slows neutrons produced by nuclear reactions so that they can be absorbed by the fuel and cause additional nuclear reactions. A coolant.

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