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Chapter 19 Acids And Bases

An acid is a substance

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that contains hydrogen and ionizes to produce hydrogen ions in aqueous solution; A base is a substance that contains a hydroxide group and dissociates to produce a hydroxide ion in aqueous solution.

Bronstred-Lowry model. A model of acids and bases in which an acid is a hydrogen-ion donor and a base is a hydrogen-ion acceptor.

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covers properties of
acids and bases, the
different definitions,
[H⁺] and [OH⁻] and pH

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and pOH, K_a and K_b
and their use when
dealing with weak
acids and bases and
buffers.

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He said that acids are
hydrogen containing
compounds that ionize
to yield hydrogen ions
(H^+) in aqueous
solution. Also said that
bases are compounds
that ionize to yield

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hydroxide ions (OH^-) in
aqueous solution

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Answer. The Arrhenius definition states that acids are compounds that contain hydrogen, and when placed in aqueous solution, ionize to form hydrogen (H^+ ions). Bases are compounds that ionize to yield hydroxide ions (OH^-) in aqueous solutions.

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SC6.

Obtain, evaluate, and communicate information about the properties that describe solutions and the nature of acids and bases. f. Use mathematics and computational thinking to compare, contrast, and evaluate the nature of acids and bases in terms of percent dissociation, hydronium ion concentration, and pH.

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Chapter 19 Acids, Bases, and Salts

Chapter 19 Acids,
Bases, and Salts 209

SECTION 19.1

ACID-BASE

THEORIES (pages
587-593) This section
compares and
contrasts acids and
bases as defined by
the theories of
Arrhenius, Brønsted-
Lowry, and Lewis. It
also identifies
conjugate acid-base

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pairs in acid-base reactions.

Chapter 19: Acids and Bases Homework Packet (50 pts)

Section 19.1 – Acid-Base Theories Acids have a sour taste, change the color of an indicator, can be strong or weak electrolytes in aqueous solution, and react with metals.

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of acid or base are added. A buffer system should contain considerably more acid than base. Specific buffer systems should be chosen based on the pH that must be maintained. 114 Study Guide for Content Mastery Chemistry: Matter and Change Chapter 19

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Acids have a sour, tart
taste, react with metals
such as Zn and Mg,
react with compounds
containing Hydroxide
ions (good proton
acceptors to form
water and a salt).

Bases have a bitter

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taste, feel slippery,
reacts with an acid to
form water and a salt,
and can form strong or
weak electrolytes in
aqueous solution. Work
Step by Step

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Arrhenius said that
acids are hydrogen-
containing compounds
that ionize to yield
hydrogen ions (H^+) in
aqueous solution. He

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also said that bases are compounds that ionize to yield hydroxide ions (OH^-) in aqueous solution.

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634 Chapter 18 • Acids
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118.18.1 Figure 18.1

Rhododendrons

flourish in rich, moist soil that is moderately acidic, whereas *semper-vivum*, commonly

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called hen and chicks,
grows best in drier,
slightly basic soil.

Introduction to Acids
and

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models help describe
the behavior of acids
and bases.

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